**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 4 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 4 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. Transition temperature of sulphur is: A) 13.2 C B) 95.5 C C)128 C D) None
2. In crystal the angle is --- between length b and c: A) alpha B) beta C) r D) S
3. Co-32 is unit: A) Tetrahedral B) Linear C) Triangle D) None
4. The carbohydrates like Glucose contain group which is responsible for hydrogen bonding in them: A) –OH B) –H C) CHO D) COOH
5. Acetone and chloroform are soluble into each other due to: A) Hydrogen bonding B) Debye Force C) London Force D) None
6. Graphite is an example of crystal system: A) Tetragonal B) Cubic C) Monoclinic D) Hexagonal
7. When water freezes H occupies % more space: A) 10 % B) 99% C) 9 % D) 1 %
8. Which of the following is pseudo solid: A) CaF2 B) Glass C) NaCl D) None
9. The polarization of Noble gases down the group: A) Decrease B) Increase C) remain Same D) None
10. Amorphous Solid: A) have sharp melting point B) undergo clean cleavage when cut with knife C)have a perfect arrangement of atom D) can posses small region of orderly arrangement of atom

**Q # 2: Short Question 10 \* 2= 20**

1. Define intermolecular forces.
2. What are dipoles induces dipole forces?
3. What do you know about London Forces?
4. Why HF is weak Acid as compare to HCl?
5. Write some uses of liquid crystal.
6. Why Ice float on water?
7. Define anisotropy.
8. Differentiate between isomorphism and polymorphism.
9. What is transition temperature?
10. What is crystal lattice?

**Q # 3 Long Questions 2\* 5 =10**

1. Explain factor effecting London Forces.
2. Define and explain seven crystal system.