**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 1 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 1 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. The phenomena of isotopes was discovered by: A)Heisenberg B) Direc C) Soddy D) Aristotle
2. At present above different isotopes occur in nature: A) 300 B) 40 C) 290 D) 280
3. The tin has ---isotopes: A) 11 B) 9 C) 8 D) 6
4. The average atomic mass of neon is: A) 20.24 amu B) 20.18 amu C) 20.19 amu d) non
5. Molecular mass of sucrose is: A) 180 g B) 280g C) 98g D) 342 g
6. Calculate the gram atom of 0.1 g of Na: A) 3.4 \* 10-3 B) 4.3 \*10-4  C) 4.9 \* 10-4  D) None
7. In 0.1 mole of H2SO4 there are total negative charge: A) 0.2 Na B) 0.1 Na C) ) 0.01 Na D) none
8. Oxygen has isotopes: A) 1 B) 2 C) 3 D) 6
9. In a gaseous state the distance between molecules is ----time greater than their diameter : A) 500 B) 200 C) 300 D) 100
10. Calculate the number of moles of 14 g of KOH: A) 0.25 B) 0.15 C) 0.75 D) 1.00

**Q # 2: Short Question 10 \* 2= 20**

1. What is stoichiometry?
2. What is actual and theoretical yield?
3. What do you know about isotopes?
4. What is Avogadro’s numbers?
5. Why actual yield is lesser than theoretical yield?
6. Calculate mole of chlorine atom in 0.22g C2 H4 Cl2.
7. Calculate the mole of Oxygen atom in 9 g of Mg (NO3)2?
8. Calculate the mass in grams of 2.74 moles of KMnO4.
9. A well known ideal gas is enclosed in in a container having volume 500 cm3 at S.T.P. Its mass come out to be 0.72 g. what is the molar mass of gas?
10. What do you know about molar volume?

**Q # 3 Long Questions 2\* 5 =10**

1. Explain limiting reactant
2. A sample of 0.600 moles of metal M reacts completely with excess of Fluorinated form 46.8 g of MF2: a) how many mole of fluorine are present in the sample of MF2 that form? B) Which element is present by symbol M?