**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 10 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 10 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. Occur at anode: A) Oxidation B) Reduction C) Both D) None
2. Of salt bridge is not used between half cell, voltage: A)decrease rapidly B) decrease slowly C) drop to zero D) None
3. Oxidation potential of Zn is: A) -0.76 B) -0.34 C) + 0.76 D) None
4. Which effect the value of reduction potential: A) Temperature B) Pressure C) Concentration D) all
5. Pure H2 gas at 1 atm is continuously bubbled into: A) 1 M HCl solution B)0.1 M HCl solution C) Both D)None
6. Potential of SHE is: A) 0.34 B) 0.76 C) 1 D) 0
7. Salt bridge contain aqueous solution of in gel: A) NaCl B) CsCl C) MgCl2 D) KCl
8. In Dawn cell anode is made of: A) Graphite B) Iron C) Titanium D) None
9. In Dawn cell is obtained by products: A) Na metal B) Cl2 C) H20 D) None
10. Oxidation number of oxygen in peroxide is: A) +2 B) -2 C) -1 D) +1

**Q # 2: Short Question 10 \* 2= 20**

1. Write rules of assigning oxidation number.
2. Write oxidation number of Cl in KClO4.
3. Differentiate between metallic conduction and electrolytic conduction.
4. Define Ionization.
5. Define electrolysis.
6. What is voltaic cell?
7. Write function of salt bridge.
8. What is SHE?
9. Define Ionization Potential.
10. What is electrochemical series?

**Q # 3 Long Questions 2\* 5 =10**

1. Write five applications of electrochemical series.
2. Explain electrolytic process of industrial important.