**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 8 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 8 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. Pure water has a pH:---- A) 7 B) 8 C) 6 D) 14
2. PKw=----@ 25C: A) 10-4 B) 10-7 C) 7 D) 14
3. If Ka < 10-3 acid is: A) strong B) very strong C) weak D) None
4. KICO3 is soluble salt: A) more soluble B) less soluble C) both D) None
5. The pH of human blood is: A) 7.35 B) 7.45 C) 7.40 D) 7.14
6. The pH of 10-3 moledm3 of aqueous solution of H2SO4 is: A) 3 B) 2.7 C) 2.0 D) 1.5
7. In contact process the crystal is: A) IV2O5 B) Pt C) both D) None
8. The low concentration of S-2 ions help to do precipitation of radical of group basic radical: A) first B) second C) third D) Fourth
9. Kc for 20 three to 30 two is 1055 @ 25 C it tell that ozone at room temperature: A) stable B) unstable C) Both D) None
10. For which system does the k is unit of (concentration) -1? A) N2 + 3H2 2NH3 B) H2 + I2 2HI C) 2 NO2 N2O4 D) 2HF H2F2

**Q # 2: Short Question 10 \* 2= 20**

1. Define law of mass action.
2. Write unit of kc for 2NH3 N2 + H3
3. What are reversible and irreversible reactions?
4. Define chemical equilibrium.
5. N2 and H2 combine to give NH3. The value of Kc in this reaction at 500 C is 0.16 \* 10 -2. Calculate the Kp for the reaction.
6. What do you know about direction of reaction?
7. Define Le-Charlie’s principal.
8. Write effect of temperature on exothermic reaction.
9. What is Ostwald dilution law?
10. Define buffer solution.

**Q # 3 Long Questions 2\* 5 =10**

1. Define synthesis of ammonia by Haber’s process.
2. Benzoic acid is weak monobasic acid (Ka = 6.4 \* 10-5 mol dm-3 ) what is pH of solution containing 7.2 g sodium benzoic in on dm-3 of 0.2 mol dm-3 benzoic acid.