**ILMI BLOG.COM FSC 1ST YEAR CHEMISTRY CHAPTER 7 TEST**

**STUDENT NAME-------------------------- ROLL # ------------------------------- DATE: / /**

**Class 1 year Chapter 7 T. Marks: 40 Subject Chemistry Time 45 min Obtain marks------------------**

**Q # 1 encircle the correct option 1\*10=10**

1. Calories is equivalent to: A) 0.4184 J B) 41.84 J C) 4.184 J C) None
2. The change in heat energy of a chemical reaction at constant temperature and pressure is called: A) enthalpy change B) Bond energy C) internal energy D) None
3. E= A) qp B) Qv C) qc D)None
4. For a given process the heat changes at constant pressure and at constant volume are related to each other: A) qp=qv B) qp>qv C) qp<qv D) None
5. For the reaction NaOH + HCl NaCl + H2O change in enthalpy is Called: A) Heat of reaction B)Heat of formation C) Heat of neutralization D) Heat of combustion
6. Stat function is a property of a system: A) Microscopic B) Macroscopic C) Both D) None
7. Heat of neutralization of NaOH + HCl? A) -14.8kjmol-1 B)+37.2 kjmol-1 C)-57.4 kjmol-1 D) None
8. C + O2 CO2 H? the heat of combustion is: A) +218kjmol-1 B) -180kjmol-1 C) -692kjmol-1 D) -393kjmol-1
9. Specific heat of water: A) 4.2 B)3.2 C) 1.2 D) 5.2
10. Pressure inside the bomb calorimeter is about: A) 30 atm B) 20 atm C) 40 atm D) 50 atm

**Q # 2: Short Question 10 \* 2= 20**

1. Define system and surrounding.
2. Define state function.
3. Define internal energy
4. Define heat
5. Define enthalpy atomization.
6. What do you know about enthalpy of solution?
7. Define Hess’s law of heat Summation.
8. Define heat of neutralization.
9. Derive enthalpy change ( H) =qp?
10. Define enthalpy.

**Q # 3 Long Questions 2\* 5 =10**

1. Explain born Haber cycle and give example of NaCl.
2. Explain Bomb calorimeter.